

More Formula Practice

Determine the formulas of the resulting compounds. (Remember to place brackets around the polyatomic ions with the subscript outside the bracket if there is more than one in the compound.)

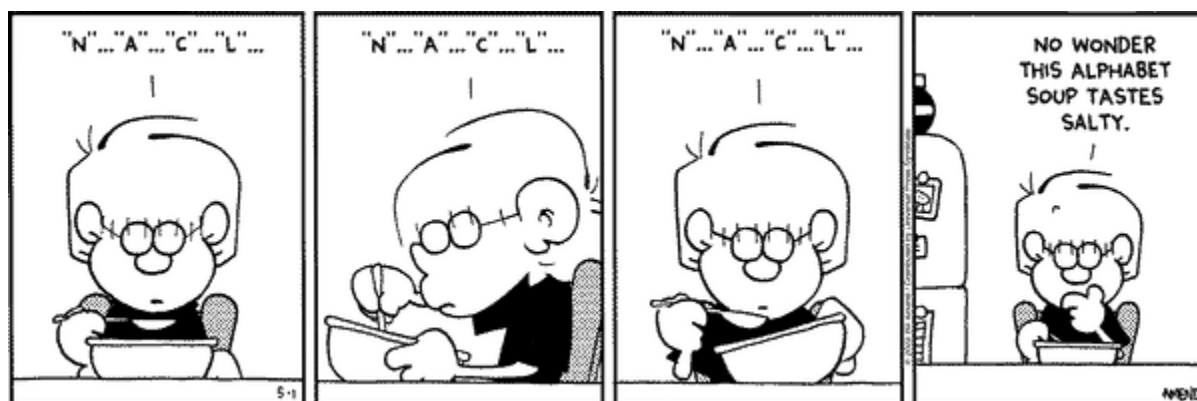
X	Cl ¹⁻	O ²⁻	P ³⁻	OH ¹⁻	CO ₃ ²⁻
Na ¹⁺					
Ca ²⁺					
Al ³⁺					
Fe ²⁺					
Fe ³⁺					
NH ₄ ¹⁺					

Note that if the names of the ions are given, it is necessary to look up the symbols and valences.

E.g. Potassium (K¹⁺) and nitrate (NO₃⁻) combine to form KNO₃.

Lithium () and sulphate () combine to form _____.

Copper (II) () and chloride () combine to form _____.



The name of an ionic compound is simply the name of the positive ion followed by the name of the negative ion, e.g.:

NaCl _____

CaCl₂ _____

Li₂O _____

KOH _____

Mg(ClO₃)₂ _____

NH₄F _____

The only tricky ones to name are those that contain multivalent ions. It is necessary to look at how the ions have combined to figure out the valence of the multivalent ion.

FeO _____
(because oxide is O²⁻)

Fe₂O₃ _____

CuCl₂ _____

CuSO₄ _____

Cu₂O _____

More Practice:

BaF₂ _____

CoCl₂ _____

AlPO₄ _____

K₃N _____

PbI₄ _____

NaOH _____

(NH₄)₂O _____

Ca₃P₂ _____

MgCO₃ _____

CuBr _____

NiCl₂ _____

KNO₂ _____

Sn₃P₄ _____

Ca₃(PO₄)₂ _____

AuCl₃ _____

SnO _____